

Unit 2 – Resources
Part A|B Exam Test Review
(20-11 2012)

Part A and B:

- Law of conservation of matter – define and apply it's meaning to a chemical (think of me burning the paper in the glass beaker)
- Balancing equations – Balance equations (q's will come from homework) (you will need to have the correct coefficients and show work) with T-table
- Chemical and physical properties – Be able to identify whether a property or change is a physical or chemical change/property
- JJ Thompson – Be able to explain his experiment with the cathode ray tube and how he discovered the electron
- Ernest Rutherford – Be able to explain how he discovered the nucleus with his experiment with alpha particles and gold foil. Be able to explain his two important conclusions.
- Dimitri Mendeleev – First periodic table 1860's
- Periodic Trends – explain with an arrow on a blank outline of the per. Table:
 - How metallic an element is
 - Ionization energy
 - Electron affinity
 - Atomic radius
- Identify: what a period is; what a group/family is
- Explain how single displacement of a metal occurs (think of the reaction of the iron nail with copper solution)
- Be able to place the metals from our lab in order of increasing reactivity (you will need to tie your results from lab to your explanation)